

Quick Naming and Formulas

Name these compounds. They may be either ionic or covalent.

- 1) LiOH Lithium Hydroxide
- 2) PBr_3 Phosphorous tri Bromide
- 3) CoSO_4 Cobalt (II) Sulfate
- 4) $(\text{NH}_4)_2\text{S}$ Ammonium Sulfide
- 5) CaCO_3 Calcium Carbonate
- 6) CF_4 Carbon tetra fluoride
- 7) CuNO_3 Copper (II) Nitrate
- 8) P_2S_3 diphosphorous trisulfide
- 9) $\text{Al}(\text{BrO}_3)_3$ Aluminum Bromate
- 10) Mg_3N_2 Magnesium Nitride

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

- 11) potassium thiocyanate KSCN
- 12) phosphorus triiodide PI_3
- 13) calcium hydroxide $\text{Ca}(\text{OH})_2$
- 14) dinitrogen sulfide N_2S_3
- 15) carbon monoxide CO
- 16) diboron tetrahydride B_2H_4
- 17) phosphorus pentabromide PBr_5
- 18) ferric chlorate $\text{Fe}(\text{ClO}_3)_3$
- 19) sodium carbonate Na_2CO_3
- 20) stannic acetate $\text{Sn}(\text{C}_2\text{H}_3\text{O}_2)_4$

Quick Atomic Structure Worksheet

Fill in the blanks for the elements in this chart. For the purposes of this chart, round all atomic masses to the nearest whole number.

Element	Number of Protons	Number of Neutrons	Number of Electrons	Atomic Mass	Atomic Number	Symbol Notation	Hyphen Notation
lithium	3	4	3	7	3	${}^7_3\text{Li}$	Lithium-7
carbon	6	6	6	12	6	${}^{12}_6\text{C}$	Carbon-12
deuterium	1	1	1	2	1	${}^2_1\text{H}$	Hydrogen-2
silver	47	61	47	108	47	${}^{108}_{47}\text{Ag}$	Silver-108
lead	82	125	82	207	82	${}^{207}_{82}\text{Pb}$	Lead-207
calcium	20	20	20	40	20	${}^{40}_{20}\text{Ca}$	Calcium-40
tantalum	73	108	73	181	73	${}^{181}_{73}\text{Ta}$	Tantalum-181